

Which of the following is FALSE about electron affinity (EA)?

1. Noble gases have the highest EAs.
2. Halogens have high EAs.
3. EA is the ability of an element or ion to gain an electron.
4. An element with a high EA is more stable as a negatively charged ion.
5. EA can be positive or negative.

Which of the following is FALSE about electron affinity (EA)?

- 72%  1. Noble gases have the highest EAs.
- 10% 2. Halogens have high EAs.
- 10% 3. EA is the ability of an element or ion to gain an electron.
- 4% 4. An element with a high EA is more stable as a negatively charged ion.
- 4% 5. EA can be positive or negative.

Which of the following statements is true?

1. An atom with a high electronegativity is an electron acceptor because it has a high affinity for electrons.
2. An atom with a high electronegativity is an electron donor because it has a high ionization energy.
3. An atom with high electronegativity can be an electron acceptor or donor because EA can be (+) or (-).
4. An atom with a high electronegativity is an electron donor because it already has too many electrons.

Which of the following statements is true?

- 88%  1. An atom with a high electronegativity is an electron acceptor because it has a high affinity for electrons.
- 6% 2. An atom with a high electronegativity is an electron donor because it has a high ionization energy.
- 5% 3. An atom with high electronegativity can be an electron acceptor or donor because EA can be (+) or (-).
- 1% 4. An atom with a high electronegativity is an electron donor because it already has too many electrons.

Group→	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	
↓Period																			
1	1 H																		2 He
2	3 Li	4 Be											5 B	6 C	7 N	8 O	9 F	10 Ne	
3	11 Na	12 Mg											13 Al	14 Si	15 P	16 S	17 Cl	18 Ar	
4	19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr	

Order the following from smallest to largest: K^+ , Na^+ , K .

1. *smallest* $\text{K}^+ < \text{Na}^+ < \text{K}$ *largest*
2. *smallest* $\text{K}^+ < \text{K} < \text{Na}^+$ *largest*
3. *smallest* $\text{Na}^+ < \text{K}^+ < \text{K}$ *largest*
4. *smallest* $\text{K} < \text{Na}^+ < \text{K}^+$ *largest*
5. *smallest* $\text{K} < \text{K}^+ < \text{Na}^+$ *largest*

Group→	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	
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4	19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr	

Order the following from smallest to largest: K^+ , Na^+ , K .

4% 1. *smallest* $\text{K}^+ < \text{Na}^+ < \text{K}$ *largest*

6% 2. *smallest* $\text{K}^+ < \text{K} < \text{Na}^+$ *largest*

85%  3. *smallest* $\text{Na}^+ < \text{K}^+ < \text{K}$ *largest*

4% 4. *smallest* $\text{K} < \text{Na}^+ < \text{K}^+$ *largest*

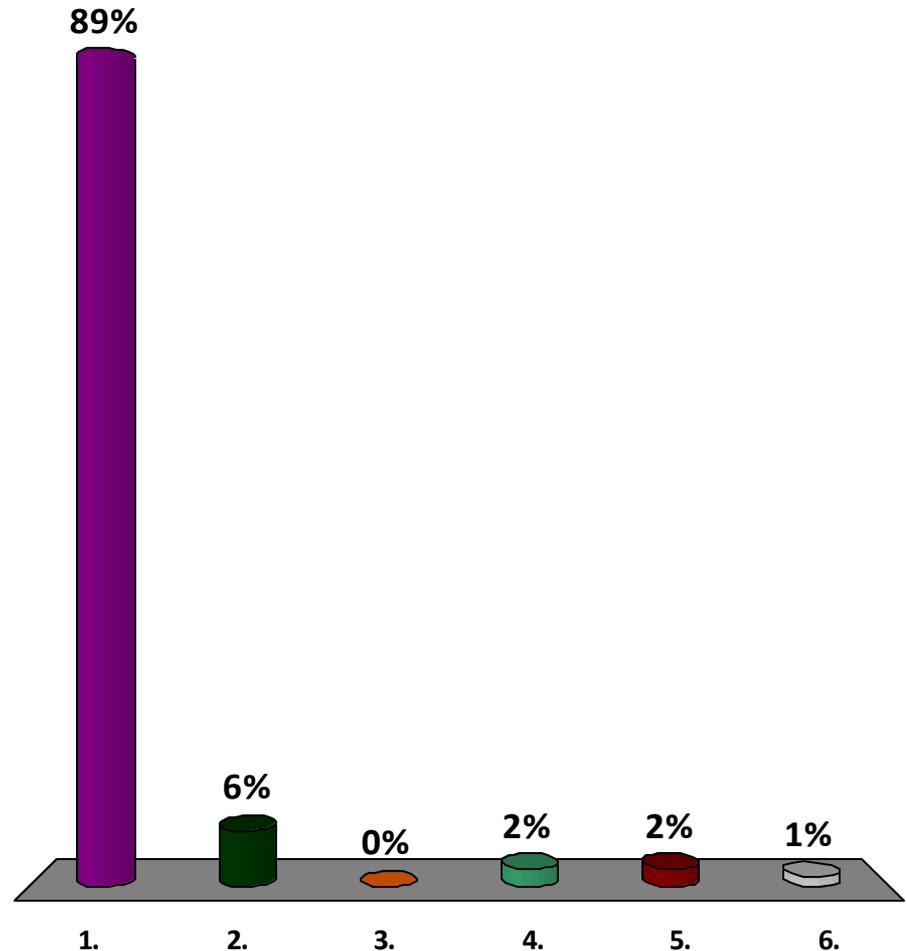
1% 5. *smallest* $\text{K} < \text{K}^+ < \text{Na}^+$ *largest*

Which of the following ions is isoelectronic with Kr ($Z = 36$)?

1. Se^{-2}
2. Se^{+2}
3. As^{-2}
4. As^{+2}
5. Rb^{-2}
6. Rb^{+2}

Which of the following ions is isoelectronic with Kr ($Z = 36$)?

- 😊 1. Se^{-2}
2. Se^{+2}
3. As^{-2}
4. As^{+2}
5. Rb^{-2}
6. Rb^{+2}



Which molecule has more polar bonds?

1. Vitamin A
2. Vitamin B9
3. Same number



Which molecule has more polar bonds?

6.86E-02 1. Vitamin A

91% 😊 2. Vitamin B9

1.96E-02 3. Same number

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